



INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

FOR THE PURPOSES OF INFORMATION ONLY

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AT	Austria	GB	United Kingdom	MR	Mauritania
AU	Australia	GE	Georgia	MW	Malawi
BB	Barbados	GN	Guinea	NE	Niger
BE	Belgium	GR	Greece	NL	Netherlands
BF	Burkina Faso	HU	Hungary	NO	Norway
BG	Bulgaria	IE	Ireland	NZ	New Zealand
BJ	Benin	IT	Italy	PL	Poland
BR	Brazil	JP	Japan	PT	Portugal
BY	Belarus	KE	Kenya	RO	Romania
CA	Canada	KG	Kyrgyzstan	RU	Russian Federation
CF	Central African Republic	KP	Democratic People's Republic of Korea	SD	Sudan
CG	Congo	KR	Republic of Korea	SE	Sweden
CH	Switzerland	KZ	Kazakhstan	SI	Slovenia
CI	Côte d'Ivoire	LI	Liechtenstein	SK	Slovakia
CM	Cameroon	LK	Sri Lanka	SN	Senegal
CN	China	LU	Luxembourg	TD	Chad
CS	Czechoslovakia	LV	Latvia	TG	Togo
CZ	Czech Republic	MC	Monaco	TJ	Tajikistan
DE	Germany	MD	Republic of Moldova	TT	Trinidad and Tobago
DK	Denmark	MG	Madagascar	UA	Ukraine
ES	Spain	ML	Mali	US	United States of America
FI	Finland	MN	Mongolia	UZ	Uzbekistan
FR	France			VN	Viet Nam
GA	Gabon				

NEW POLYMERIC CONDUCTIVE ALUMINO-SILICATE MATERIAL, ELEMENT
COMPRISING SAID MATERIAL, AND PROCESS FOR PREPARING IT.

The present invention concerns a novel polymeric
5 conductive alumino-silicate and a process for preparing it.
The present invention also concerns an element having thereon
said new polymeric material.

Aluminosilicate in fiber form is known in the Art. A
filamentous, tubular and crystallized aluminosilicate called
10 Imogolite has been found in volcanic ashes and other soils.
Farmer in USP 4,152,404 and 4,252,779 describes an inorganic
material related to natural imogolite. This inorganic
material is synthesized by reacting soluble silica or a
soluble silicate compound with an aluminum compound to form a
15 soluble hydroxyaluminum silicate complex in an aqueous
solution at a pH of 3.2 to 5.5 ; and digesting at pH 3.1 to
5.0 soluble hydroxyaluminum complexes freshly formed, thereby
forming a colloidal dispersion of the inorganic material.
This inorganic material can be used as molecular sieve,
20 catalyst support, coagulant or sorbent. By evaporating
imogolite colloidal solutions onto a flat surface, coherent
films can be formed and used as membranes. When the product
is not isolated from its colloidal solution, it may be used
as a flocculant, a hydrophilizer or a thickener.

25 Aluminum silicon metal oxide particles have been
described in a large number of patents. These particles are
crystallized electrically conductive particles which are used
dispersed in a binder as antistatic material. For example, a
composite oxide from aluminum and silicon has been described
30 in US 4,495,276. In this patent, the crystalline metal oxide
particles are used as an antistatic layer in a photographic
element.

EP 250,154 describes a photographic element comprising a
polymeric surface to which is adhered a layer comprising a
35 gelled network of inorganic particles, preferably inorganic
oxide particles. According to that patent, the gelled

network is a porous coating having voids between the inorganic oxide particles. The porous coating is obtained from a dispersion or suspension of finely divided particles in a liquid medium. Inorganic oxide particles can be 5 boehmite (aluminum oxide), silica or alumina-coated silica. The "gelled network" is formed by an aggregation of colloidal particles linked together to form a porous three-dimensional network. The gelled network provides a subbing layer having antistatic properties.

10 An object of the present invention is to provide a new homogeneous polymeric alumino-silicate material having antistatic properties. An other object of the present invention is to provide a new process to obtain this polymeric material. A third object of the present invention 15 is to provide an element comprising an antistatic layer consisting of the polymeric material of the present invention.

In accordance with the invention, there is provided a process to obtain the polymeric alumino-silicate according to 20 the invention which comprises the following steps :

(a) treating a mixed aluminum-silicon alkoxide, or a mixed aluminum-silicon precursor, with an aqueous alkali, whilst maintaining the pH at a value between 4 and 6.5 ; the Al molar concentration being maintained in the range of 5×10^{-4} 25 to 10^{-2} , and the Al/Si molar ratio between 1 and 3 ;
(b) heating the mixture obtained in step (a) at a temperature below 100°C, and in the presence of silanol groups, for a period sufficient to have a complete reaction whereby a polymeric material is obtained ; and
30 (c) eliminating the residual ions from the reaction medium.

The material thus obtained is a polymeric fibrous alumino-silicate which can be directly applied onto a support in order to unexpectedly form an antistatic layer.

An other object of the present invention is a polymeric 35 inorganic material comprising more than 80 moles % of a fibrous alumino-silicate of formula $\text{Al}_x\text{Si}_y\text{O}_z$, wherein the

ratio x:y is between 1 and 3, preferably between 1.5 and 2.5 ; z is between 1 and 10, preferably between 2 and 6. According to one embodiment, the polymeric material comprises more than 98 moles % of a fibrous alumino-silicate having the 5 above formula. This new polymeric material has properties which make it possible to envisage its use for producing in particular antistatic coatings or permeable membranes.

In accordance with the invention, it is also provided an element comprising a support having thereon an antistatic 10 layer consisting of a polymeric inorganic material based on silicon and aluminum comprising more than 80 moles % of a fibrous alumino-silicate of formula $Al_xSi_yO_z$ in which the ratio x:y is between 1 and 3, and z is between 2 and 6.

By the present invention, it is further provided a 15 photographic element comprising a support having thereon at least one silver halide emulsion and an antistatic layer obtained by coating on the support the material obtained after eliminating the residual ions.

The novel polymeric electroconductive alumino-silicate 20 has a fractal morphology demonstrated by transmission electron microscopy, a homogenous fibrous structure demonstrated by electron diffraction and a composition characterized by a homogenous Al/Si molar ratio demonstrated by X-ray microanalysis (using an energy dispersive X-ray 25 spectrometer).

Figure 1A is a transmission electron photomicrograph of the material of the present invention. Figure 1B is a transmission electron photomicrograph of a comparative material.

30 Figure 2A is a transmission electron photomicrograph of the material of the present invention. Figure 2B is a transmission electron micrograph of the comparative material.

Figure 3 is a transmission electron photomicrograph of the material obtained from the Farmer's process disclosed in 35 USP 4,252,779.

An embodiment of the above process consists of forming

in situ a mixed aluminum and silicon precursor by reacting an aluminum salt and a silicon alkoxide. The aluminum salt may be a halide, for example a chloride or bromide, a perhalogenate (for example perchlorate), a sulphate or a nitrate, a phosphate or an aluminum salt of an organic acid. Preferably, a halide or a perhalogenate is used. The silicon alkoxide can be a methoxide or an ethoxide, or a corresponding chloroalkoxide. The hydrolysis of the aluminum salt and the silicon alkoxide is effected in a controlled manner, in order to form the mixed aluminum and silicon precursor which can then be used in step (a) defined above.

The controlled hydrolysis is carried out by slowly adding the silicon alkoxide to an aqueous solution of the freshly prepared aluminum salt. In practice, the silicon alkoxide, either pure or diluted in an alcohol solution, is added dropwise to an aqueous solution of aluminum chloride, aluminum perchlorate, etc. The addition is effected at room temperature, with stirring, the aluminum solution having a concentration of between 5×10^{-4} M and 10^{-2} M and advantageously between 5×10^{-3} M and 7×10^{-3} M. This technique is known as the sol/gel method, described for example in French patent application 9203653 or in US patent 4,122,041. This controlled hydrolysis can also be carried out using the so-called "inverted micelles" technique, described in European patent EP 332,686.

A mixed aluminum-silicon precursor can also be formed by mixing a silicon alkoxide or chloroalkoxide with an aluminum alkoxide or chloroalkoxide, for example an alkoxide group having from 1 to 5 carbon atoms such as a methoxide, a propoxide or isopropoxide, or a butoxide, and then hydrolyzing this mixture, as described above.

The process according to the invention comprises a first step (a) which consists of treating either the mixed aluminum-silicon alkoxide or the mixed aluminum-silicon precursor by means of an alkali in aqueous or hydroalcoholic solution. Step (a) should be carried out whilst maintaining

the pH between 4 and 6.5 and preferably between 4.5 and 5.5; the optimum pH depending on the starting compounds used. If the pH is below 4, the action of the alkali results in the non-reversible formation of hydrated alumina, whilst at a pH 5 above 6.5 there is the non-reversible formation of a silica gel.

The alkali is preferably an aqueous solution of sodium or potassium hydroxide, at a concentration of between 2 M and 10^{-2} M, preferably between 1 M and 10^{-2} M. The alkali 10 solution is added to the mixed aluminum-silicon alkoxide or to the mixed aluminum-silicon precursor at a rate of between 0.1 and 100 mmoles/hour and advantageously between 1 and 10 mmoles/hour. The total quantity of alkali added represents, in moles, between 2 and 3 times and preferably between 2.3 15 and 2.5 times the amount of aluminum.

In a second step (b), the process according to the invention consists of heating the mixture resulting from the previous step (a) under predetermined temperature conditions, in the presence of silanol groups, until a complete 20 condensation reaction is achieved. By "condensation reaction, it is meant "a reaction whereby the alkoxide molecules are co-condensed with elimination of H_2O ". The condensation reaction is considered to be completed when the reacting medium contains no ions other than those coming from 25 the alkali, that is to say Na^+ or K^+ ions; all the other ions, that is to say those of the silicon and aluminum, having been used to form a polymeric aluminosilicate.

Alternatively, the mixture resulting from step (a) can be concentrated, submitted to a treatment such as 30 centrifugation to isolate a gel and diluted again before heating step (b). Such a procedure is illustrated by example 1 herein below.

According to one embodiment, step (b) is carried out at a temperature in the range of from about 70 to 98°C.

35 An important characteristic of the invention is that this heating step is carried out in the presence of silanol

groups.

The tests carried out within the framework of the present invention show particularly that the condensation reaction in the above step (b) produces an inorganic polymeric alumino-silicate with a homogeneous fibrous structure, provided that the reaction is carried out in the presence of silanol groups. For example, the silanol groups can be introduced by means of a substrate having an amount of accessible silanol (SiOH) groups, that is to say located on the surface of the substrate, which is at least approximately 4 silanol groups per μm^2 of surface area of substrate, which corresponds to a ratio between the number of Al atoms engaged in the reaction to the number of silanol groups of less than approximately 2.25×10^{10} and advantageously less than 1.0×10^{10} .

If the process is carried out in a glass reactor, the silanol groups are provided by the internal wall of the reactor itself as Example 1.

Preferably, the substrate comprising superficial SiOH silanol groups is in a divided form. An appropriate implementation of the process of the invention consists of carrying out step (b) in an inert reactor, this is to say one with no silanol radical on its internal surface and/or not including any ancillary device such as a stirrer or probe with such silanol groups, and adding the desired amount of silanol groups to the reaction medium. In this way, it is possible to control this step of the process. The inert reactor is for example a reactor made from stainless steel or a polymeric material such as polytetrafluoroethylene; it is possible to introduce the appropriate substrate with silanol groups, for example using various plates pre-treated in order to graft onto the surface the desired quantity of accessible silanol groups, and then immersed in the reaction medium. The number of SiOH equivalents required for carrying out the reaction in order to obtain the intended product was determined in accordance with the information provided by

B. Arkles in Silicon Compounds; Register and Review, Petrarch Systems Silanes & Silicones, 1987, ABCR GmbH & Co, Karlsruhe, page 54.

During the heating step (b), the Al molar concentration
5 is in the range of from 5×10^{-4} to 10^{-2} M, and preferably
 2.5×10^{-3} to 9×10^{-3} M. During the heating step (b), the
pH value is preferably in the range from 4 to 5, but can be
varied depending on the selection of the starting reactants.

Step (c) of the present invention consists in
10 eliminating the residual ions which are present in the
reaction medium at the end of step (b). The ions designated
as residual ions are essentially the ions from the alkali
used in step (a). The elimination of the residual ions can
be carried out for example by dialysis or by ultra-
15 filtration.

The Al:Si molar ratio, calculated with regard to the
starting products, is between 1 and 3 and preferably between
1.5 and 2.5. In the final material, the Al:Si molar ratio is
substantially the same.

20 The polymeric material resulting from step (c) can be
isolated from its solution by lyophilisation, or it can be
concentrated by gelling and centrifugation. In the isolated
form, the material can be used as a membrane or molecular
sieve, for filtration applications. The gel can be applied
25 directly onto a support or the colloidal solutions of the
polymeric alumino-silicate can also be evaporated in order to
produce deposits on various surfaces. The coatings obtained
have antistatic properties which are stable over time.

Figure 1A shows an example of a view obtained by
30 transmission electron photomicrograph. This figure shows the
morphology of the material. This morphology is observed at
increasing definitions/magnifications which enables it to be
classified as fractal. This morphology and fibrous structure
are also observed whatever the location of the examination:
35 this structure is the same at any point on the material
(magnification: 100,000).

Figure 1B shows an example of a transmission electron photomicrograph obtained with a comparative material. The heterogeneity of the structure can be seen.

Figure 2A shows an example of a transmission electron photomicrograph obtained with a sample of the material according to the invention, with a higher magnification than for Figure 1A.

Figure 2B shows an example of a transmission electron photomicrograph obtained with a comparative sample, with a greater magnification than for Figure 1B.

Figure 3 is a transmission electron photomicrograph obtained with the Farmer's material prepared according to the process of example 4. This figure clearly shows that the obtained material presents a heterogeneous structure.

The process according to the invention makes it possible to obtain, directly and without any purification operation, a synthetic inorganic material based on aluminum and silicon, characterized in that it consists, for more than 80 moles %, and advantageously more than 90 moles % and, even more advantageously, more than 98 moles %, of a fibrous alumino-silicate of formula $\text{Al}_x\text{Si}_y\text{O}_z$, where x, y and z have the meanings given above. The present material is unexpectedly homogeneous.

When the operating conditions of the process according to the invention are not used, a heterogeneous product is obtained, comprising up to 60% molar alumina (boehmite) in the form of aggregates. Such a product, for which the characterization views are given in Figs 1B and 2B does not make it possible to obtain stable gels and consequently does not lend itself to the production of either antistatic or materials for filtration. Furthermore, the conductivity of the material of the present invention is particularly stable with time.

The following examples illustrate the invention.

EXAMPLE 1

16.7 mmoles of tetraethylorthosilicate, Si(OEt)_4 , was added to 1000 ml of deionized water. The reaction mixture was stirred at room temperature for 1 hour, and then this 5 solution was added to a solution of 31.2 mmoles of aluminum chloride hexahydrate $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ in 1000 ml of pure water.

The mixture was stirred vigorously for 20 minutes, and then a NaOH, 1M solution was added to the colorless solution until a pH of 4.5 was obtained. A cloudy solution was 10 obtained, which was stirred overnight.

The pH was adjusted to 6.8 with NaOH, 1M. A white gel was obtained which is centrifuged for 20 minutes at 2000 rpm. A white gel was collected which was solubilized with 5 ml of a 50:50 mixture of hydrochloric acid 1M and acetic acid 2M. 15 The volume of the solution was made up to 1 liter with water. The solution contains 30 mmoles of aluminum, 16.6 mmoles of silicon, 5 mmoles hydrochloric acid and 10 mmoles acetic acid. This solution, as referred as the mother liquor was kept at 5°C for subsequent use.

20 250 ml of the mother liquor obtained above were diluted in 550 ml of deionized water and the mixture was heated at a temperature of between 95 and 100°C for 2 hours in a 20 cm diameter glass spherical reactor. The quantity of SiOH groups on the surface of the internal wall of the reactor, in 25 contact with the reaction medium, was estimated not less than 4×10^{11} .

The heating step was carried out as follows : every two days, 250 ml of the mother solution was added to the solution heated in the presence of silanol groups. After 120 hours of 30 heat treatment, the solution was cooled down at room temperature. At this solution, an ammonia solution NH_4OH , 1M was added slowly to adjust the pH to 8.0. A gel was obtained. The gel obtained was centrifuged for 15 minutes at 2000 rpm and the supernatant solution was removed.

35 The remaining gel was solubilized by the addition of a few drops of hydrochloric acid 12N, and then dialysis was

10

carried out with deionized water for four days. A thixotropic gel was obtained.

This gel, by lyophilisation, provided a cotton-like fibrous material, which can be stored without deterioration 5 for an indefinite period.

A solution of 10 g of the above gel was used. An Estar® polyester support was coated with this gel in order to obtain approximately 30 mg of inorganic polymer per m². This layer has a resistivity of 10⁸ ohms/square at 25 % RH, 22°C, stable 10 over time which makes it possible to envisage its use as an antistatic material.

EXAMPLE 2

In a 20 cm diameter spherical reactor made from inert 15 plastic (polytetrafluoroethylene), 2.4 g (10 mmoles) of AlCl₃.6H₂O, purity 99%, sold by the Company Aldrich, were mixed with 2000 ml of deionized water.

1.05 g (5 mmoles) of Si(OEt)₄, purity 99%, sold by the Company Ventron, was added using a syringe and with vigorous 20 stirring. The mixture was left to stand for 20 minutes. The Al/Si molar ratio was 2.

220 ml of NaOH, 0.1 M were added slowly whilst stirring. The mixture was left to stand for 2 hours. The pH was then 4.35 at 25°C.

25 A first sample A of 900 g of this solution was introduced into an inert plastic flask containing 50 g of glass chips (Raschig rings), and a second sample B of 1000 g of the same solution, was introduced into an identical flask but without adding the glass chips..

30 The two samples A and B were heated in an oven at 98°C for 5 days.

The two samples were then gelled by the addition of ammonia solution and then centrifuged in order to separate the gel from supernatant.

35 The following were recovered:

318.78 g of gel with sample A

370.75 g of gel with sample B.

Each gel was separately redispersed in HCl, 12N, and then dialyzed in 17.2 l of deionized water. 310.08 g of gel was recovered with sample A (31%) and 304.02 g with sample B
5 (40%). Samples A and B were examined by transmission electron microscopy. Figure 1B shows clearly a heterogeneous structure with particles, whilst Figure 1A (sample A according to the invention, prepared in the presence of silanol) shows an uniform fibrous structure. Figures 2A and
10 2B show respectively the same materials under greater magnification, which confirm the observations made with regard to Figures 1A and 1B.

The samples A (invention) and B (comparison) were analyzed by EDXRMA (energy Dispersive X ray Micro Analyser).

15 The Sample B spectrum shows a high Aluminium to silicon peak height ratio corresponding to a molar ratio greater than 4. This ratio is consistent with the boehmite structure dominating the reaction products and absorption of silica species. Analyses performed using a nanoprobe on various
20 areas gives a wide range of Al/Si ratios reflecting variations of the amount of adsorbed silica species.

The sample A spectrum shows a corresponding Al/Si peak height ratio estimated to 1.60 to 1.80 which is in a good agreement with the Al/Si molar optimum ratio found for the
25 formation of Imogolite. This apparent concentration has been obtained with various probe sizes (2.5 μm to 2 nm) illustrating the remarkable chemical homogeneity of this material.

The samples A (invention) and B (comparison) were also
30 analyzed by Electron Diffraction. From Sample A, a highly ordered pattern typical of fibrous structure was obtained whereas the pattern obtained from sample B corresponds to a non fibrous structure.

35 EXAMPLE 3

The procedure of example 1 was repeated except that, for

12

the heating step all the mother liquor was used in a single batch. Then, after the heating step was completed, the resulting product was filtered through an ultra filtration membrane having a molecular weight cut-off of 20,000
5 manufactured by Osmonics® corp., with a flow rate of 550 ml/min. A clear and high viscosity solution was obtained having an Al/Si ratio of 1.8.

EXAMPLE 4 : Reproduction of the process as described in
10 US 4,252,779
25.8 10^{-3} mol. of $\text{Na}_2\text{SiO}_3 \cdot 5\text{H}_2\text{O}$ was dissolved in
1614 cm^3 of deionized water and 53 cm^3 of Na_2CO_3 , 1M in order
to obtain a first solution.
15 50 mmol. of $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ was mixed with 291 cm^3 of
deionized H_2O and 42 cm^3 of HClO_4 , 1M in order to obtain the
second solution.

After mixing the first and the second solution, pH was
adjusted at 4.5 with NaOH , 1M. When the solution was clear,
pH was adjusted at 6.8 with NaOH , 1M. A precipitate was
20 obtained which was recovered by centrifugation (3000 rpm).
The recovered precipitate was solubilized with 15 cm^3 of a
solution containing HCl , 1M and CH_3COOH , 2M. The volume was
adjusted at 1000 cm^3 with water.

200 cm^3 of the obtained solution were diluted with
25 800 cm^3 of deionized water in order to obtain a Al molar
concentration 10 mmol/l. The diluted solution was maintained
at 96°C for 5 days, in a plastic flask. At room temperature,
the pH was adjusted at 8,00 by introduction of NH_4OH .
456.83 g of gel were recovered after centrifugation at
30 3000 rpm for 30 min. The gel was solubilized with HCl , 12N,
and dialized with deionized water. The gel thus obtained was
analyzed according to the method disclosed above.

The EDXRMA spectrum of this material is identical to the
spectrum of sample A (Imogolite structure), however the
35 transmission electron photomicrograph (Figure 4) shows that
the obtained material presents a heterogeneous structure.

EXAMPLE 5

In Example 5.1, the gel of Example 1 obtained after elimination of the residual ions (invention) was coated on a support. In example 5.2, the gel of Example 4 was coated on a support. The resistivity of both layers thus obtained was measured immediately after the coating, after 6 days, and after 1 month.

The results are summarized in following Table 1.

10

TABLE 1

	Resistivity (Ω/square) 20 % RH, T=22 °C			Coverage (Al+Si) (mg/m ²)
	Fresh layer	After 6 days	After 1 month	
Ex. 5.1 (INV)	7×10^7	1×10^8	1×10^8	30
Ex. 5.2 (comp.)	2.45×10^7	3×10^9	$> 10^{12}$	30

It is clear from these results that the antistatic properties of the layer of the present invention, e.i., a layer consisting essentially of fibrous alumino-silicate of the present invention are stable with time. The high resistivity of the fresh layer of Example 5.2 is not due to the material itself but it is due to the presence of water into the coated material. This is not the case with the material of the present invention.

EXAMPLE 6

6.1 - The material obtained in example 1 after elimination of the residual ions was coated as antistatic layer (30 mg/m² based on Al+Si) on one side of a support, the support having on the other side a silver halide emulsion layer and an overcoat. The resistivity of the photographic element was $7 \times 10^7 \Omega/\text{square}$. This resistivity was stable with time.

The sensitometric results obtained with and without the antistatic layer show that the presence of the antistatic layer does not modify the sensitometric properties of the photographic element.

5

6.2 - The same antistatic layer was coated on the overcoat of the photographic element as above described. The resistivity of the photographic element was about $10^9 \Omega/\text{square}$. This resistivity was stable with time.

10 As in the previous case, no modification of the sensitometric properties was noticed.

In all the examples, the electron photomicrographs were
15 produced with a Philips CM12-STEM microscope, used in transmission mode. The microscope was calibrated with a grid of 2160 lines/mm in the magnification range $\times 30,000 - \times 160,000$. Examination of the samples was carried out at a voltage of 120 kV. The X-ray diffraction spectra were
20 produced with an EDAX Micro Analyzer connected to the CM12 microscope. The calibration was carried out with samples of silica and kaolin.

CLAIMS

1. A process for preparing a polymeric inorganic conductive material based on aluminum and silicon, comprising
5 the following steps:

(a) treating a mixed aluminum-silicon alkoxide, or a mixed aluminum-silicon precursor, with an aqueous alkali, whilst maintaining the pH at a value between 4 and 6.5 ; the Al molar concentration being maintained in the range of
10 5×10^{-4} to 10^{-2} , and the Al/Si molar ratio between 1 and 3;

15 (b) heating the mixture obtained at step (a) at a temperature below the boiling point of water, in the presence of silanol groups, for a period sufficient to have a complete reaction whereby a polymeric material is obtained; and

(c) eliminating the residual ions from the reacting medium.

2. The process of Claim 1, wherein, in step (b), the amount of silanol groups corresponds to a ratio between
20 aluminum atoms and silanol groups of less than 2.25×10^{10} .

3. The process of Claim 1, wherein, in step (b), the amount of silanol groups corresponds to a ratio between aluminum atoms and silanol groups of less than 1.0×10^{10} .

4. The process of Claim 1 wherein the Al/Si molar
25 ratio is between 1.5 to 2.5.

5. The process of Claim 1, wherein, in step (b), said mixture is heated at a temperature in the range from about 70 and 98°C.

30 6. The process of Claim 1, wherein, said alkali is selected from the group consisting of sodium or potassium hydroxide.

7. The process of Claim 6 , wherein said alkali is added at a rate of 0.1 to 100 mmoles per hour.

35 8. The process of Claim 7, wherein said alkali is added at a rate of 1 to 10 mmoles per hour.

9. The process of Claim 1 wherein, in step (a),

the pH is maintained at between 4.5 and 5.5.

10. The process of Claim 1 to 9, wherein said mixed precursor compound of aluminum and silicon is the product of the hydrolysis of (i) a compound selected from the 5 group consisting of an aluminum salt, an aluminum alkoxide and an aluminum halogenoalkoxide and (ii) a silicon alkoxide or silicon chloroalkoxide.

11. The process of Claim 10, wherein the mixed precursor compound of aluminum and silicon is the product of 10 the hydrolysis of an aluminum salt or an aluminum alkoxide or halogenoalkoxide and a silicon alkoxide.

12. The process of Claim 11, wherein said aluminum salt is selected from the group consisting of a nitrate, a phosphate, a sulphate or the aluminum salt of an organic 15 acid.

13. The process of Claim 11, wherein said silicon alkoxide is an ethoxide or a methoxide.

14. The process of Claim 1 wherein, the mixed aluminum-silicon precursor as obtained by slowly introducing 20 a silicon alkoxide into a freshly prepared aluminum salt aqueous solution.

15. A polymeric inorganic material based on silicon and aluminum comprising more than 80 moles % of a fibrous alumino-silicate of formula $\text{Al}_x\text{Si}_y\text{O}_z$ in which the 25 ratio x:y is between 1 and 3, and z is between 2 and 6.

16. A polymeric inorganic material according to claim 15 comprising more than 98 mol. % of a fibrous alumino-silicate.

17. An element comprising a support having thereon 30 a conductive layer consisting of a polymeric inorganic material based on silicon and aluminum comprising more than 80 moles % of a fibrous alumino-silicate of formula $\text{Al}_x\text{Si}_y\text{O}_z$ in which the ratio x:y is between 1 and 3, and z is between 2 and 6.

35 18. A photographic element comprising a support having thereon at least one silver halide emulsion and a

17

antistatic layer obtained by coating on the support the material obtained by the process of any of preceding claims 1 to 14.

1/5

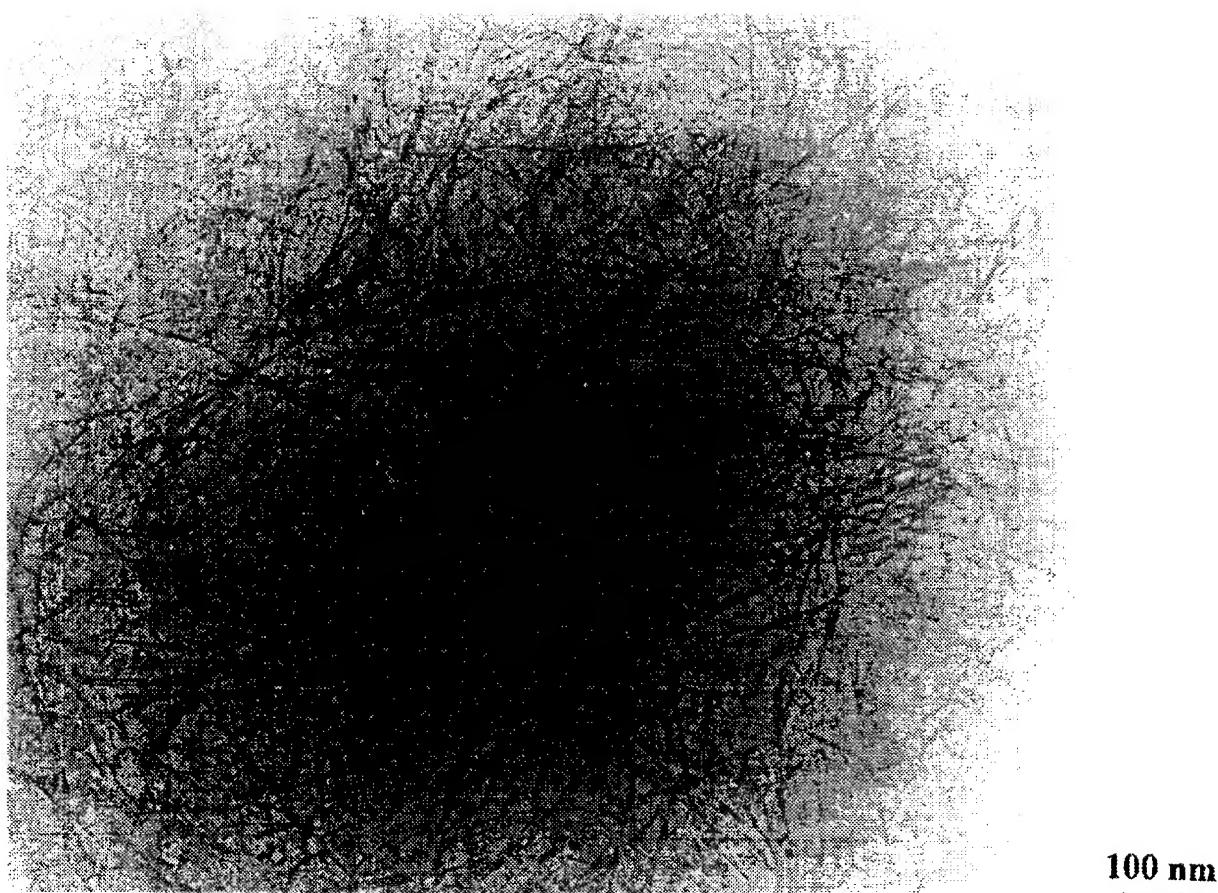


Fig. 1A

2/5



Fig. 1B

3/5

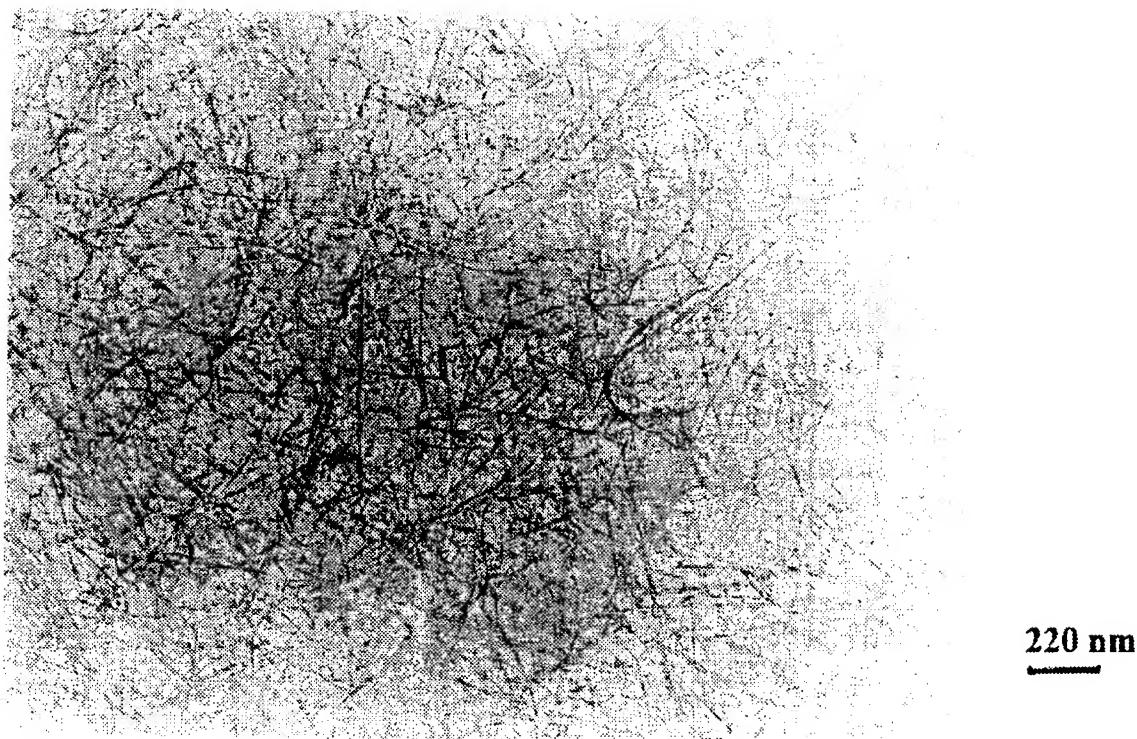


Fig. 2A

4/5



Fig. 2B

5/5

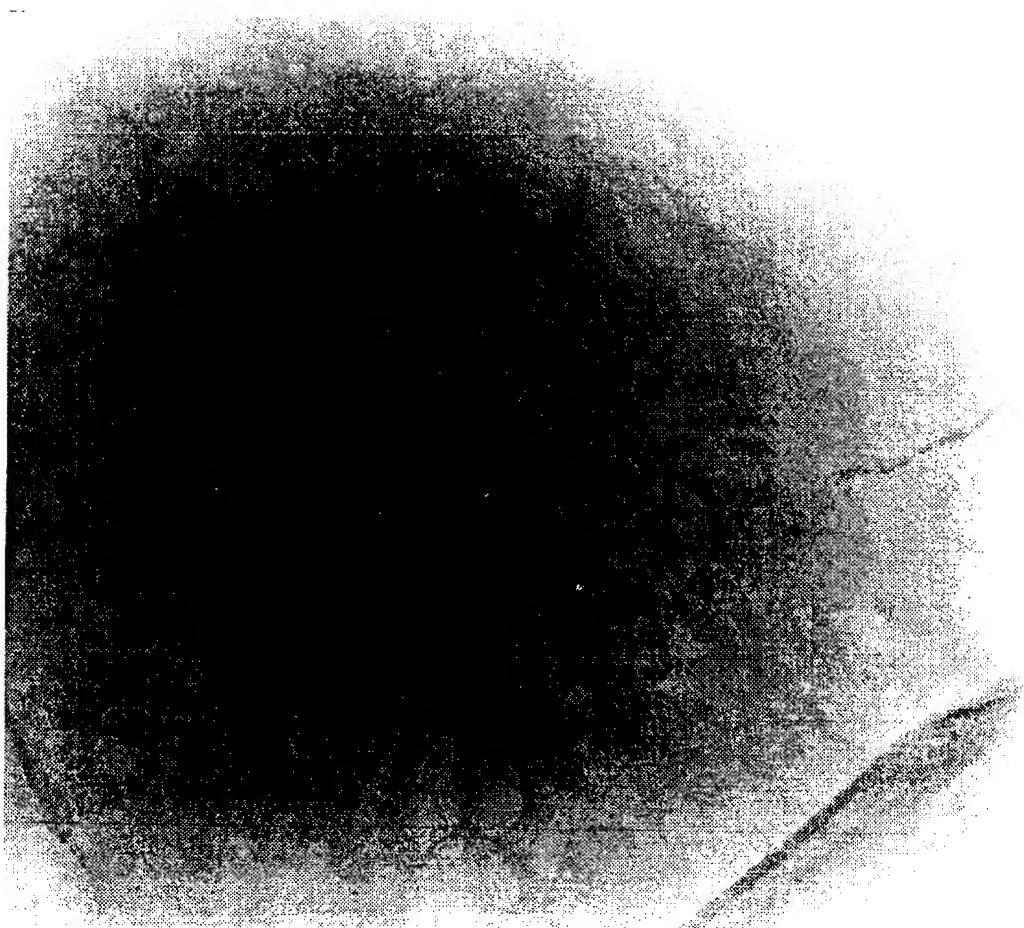


Fig. 3

INTERNATIONAL SEARCH REPORT

International Application No

PCT/EP 95/04165

A. CLASSIFICATION OF SUBJECT MATTER
 IPC 6 C01B33/26 H01B1/08 G03C1/85

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
 IPC 6 C01B G03C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	GB,A,2 025 384 (THE MACAULAY INSTITUTE FOR SOIL RESEARCH) 23 January 1980 see claims 1-24 see page 1, line 6 - line 34 ---	1,15,16
A	CHEMICAL ABSTRACTS, vol. 91, no. 22, 26 November 1979 Columbus, Ohio, US; abstract no. 178284, WADA S. I. & AL. 'Synthetic allophane and imogolite' see abstract & J. SOIL SCI., vol. 30, no. 2, 1979 pages 347-355, --- ---	1,15,16

 Further documents are listed in the continuation of box C. Patent family members are listed in annex.

* Special categories of cited documents :

- 'A' document defining the general state of the art which is not considered to be of particular relevance
- 'E' earlier document but published on or after the international filing date
- 'L' document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- 'O' document referring to an oral disclosure, use, exhibition or other means
- 'P' document published prior to the international filing date but later than the priority date claimed

- 'T' later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- 'X' document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- 'Y' document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- '&' document member of the same patent family

1

Date of the actual completion of the international search 30 January 1996	Date of mailing of the international search report 16.02.96
Name and mailing address of the ISA European Patent Office, P.B. 5818 Patentaan 2 NL - 2280 HV Rijswijk Tel. (+ 31-70) 340-2040, Tx. 31 651 epo nl, Fax (+ 31-70) 340-3016	Authorized officer Rigondaud, B

INTERNATIONAL SEARCH REPORT

International Application No

PCT/EP 95/04165

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	EUROPEAN POLYMER JOURNAL, vol. 27, no. 7, 1991 OXFORD GB, pages 609-612, BARRETT S. M. & AL. 'The synthesis and characterization of imogolite' * page 609: 'Experimental procedures; Method I' * * pages 610,611: 'Results and discussion; Synthetic method I' * ---	1,15,16
A	FR,A,2 385 646 (NATIONAL RESEARCH DEVELOPMENT CORPORATION) 27 October 1978 see claims 1,2 ---	1,15,16
A	GB,A,2 092 768 (FUJI PHOTO FILM CO. LTD.) 18 August 1982 see claim 1 see page 1, line 61 - page 2, line 11 & US,A,4 495 276 cited in the application ---	17,18
A	EP,A,0 511 764 (KONICA CORPORATION) 4 November 1992 see claim 1 ---	17,18
A	DATABASE WPI Section Ch, Week 9432 Derwent Publications Ltd., London, GB; Class A89, AN 94-257734 & JP,A,06 186 672 (KONICA CORP) , 8 July 1994 see abstract -----	18

INTERNATIONAL SEARCH REPORT

International Application No PCT/EP 95/04165

Patent document cited in search report	Publication date	Patent family member(s)		Publication date
GB-A-2025384	23-01-80	EP-A, B US-A-	0007713 4241035	06-02-80 23-12-80
FR-A-2385646	27-10-78	GB-A- BE-A- CA-A- CH-A- DE-A- JP-C- JP-A- JP-B- NL-A- US-A- US-A-	1574954 865317 1094778 641127 2813323 1441904 53124199 62038283 7803281 4252779 4152404	10-09-80 17-07-78 03-02-81 15-02-84 05-10-78 30-05-88 30-10-78 17-08-87 02-10-78 24-02-81 01-05-79
GB-A-2092768	18-08-82	JP-C- JP-A- JP-B- DE-A- US-A-	1320712 57104931 60049894 3150514 4418141	11-06-86 30-06-82 05-11-85 08-07-82 29-11-83
EP-A-511764	04-11-92	JP-A- US-A-	4328741 5294525	17-11-92 15-03-94